Tetrahedron Letters, Vol.26, No.11, p 1362, 1985
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# ERRATA

"Selective reduction of aromatic nitro compounds with stannous chloride in non acidic and non aqueous medium"

F.D. Bellamy and K. Ou
Tetrahedron Letters, Vol.25, No.8, pp 839-842, 1984.

The authors would like to publish the corrected procedure below:

## REDUCTION OF P-NITROBENZOIC ACID

A mixture (6)of 1.67g (0.01 mole) of p-nitrobenzoic acid and 11.275 g (0.05 mole) of $SnCl_2$, $2H_2O$ (7) in 20 ml of absolute ethanol is heated at 70°C under nitrogen. After 30 mn the starting material has disappeared and the solution is allowed to cool down and then poured into ice. The pH is made slightly basic (pH 8) by addition of 5 % aqueous sodium bicarbonate (8) and the resulting basic mixture kept one hour under stirring[+]. Addition of acetic acid allows the pH to decrease to 5[++]. The aqueous mixture is extracted three times with ethyl acetate, the organic phase thoroughly washed with brine, treated with charcoal and dried over sodium sulphate. Evaporation of the solvent leaves 1,295 g (94.5 %) of p-aminobenzoic acid, which gives one spot on tlc and melts at 187°C.

[+] When 5 % aqueous sodium bicarbonate is used for neutralisation, it is essential to keep the mixture one hour under stirring in order to allow the tin salts to hydrolyse. This step can be speeded up by using aqueous NaOH and a more basic pH ( $\veebar$ 10).

[++] Other amino compounds can be extracted directly from the basic aqueous phase without acidification. Only in the case of p-aminobenzoic acid is it necessary to free the acid from its salt by making the pH slightly acid.

We would like to express our gratitude to Dr. R.G. PLEVEY at Birmingham University for having called our attention on the uncorrectness of the published procedure.

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"Umpaired electron distribution in acylaminyls"

A.R. Forrester, H. Irikawa and G. Soutter
Tetrahedron Letters, Vol.25, No.47, pp 5445-5448, 1984.

Author G. Soutter - spelling corrected.